TOROS ÜNİVERSİTESİ

Vocatıonal School
Constructıon Technology

Course Information

CHEMISTRY					
Code	**Semester**	**Theoretical**	**Practice**	**National Credit**	**ECTS Credit**
		Hour / Week			
ITP109	Fall	2	0	2	2

Prerequisites and co-requisites	NONE
Language of instruction	Turkish
Type	Required
Level of Course	Associate
Lecturer	Lect. Dilşat AKGÜL
Mode of Delivery	Face to Face
Suggested Subject	NONE
Professional practise (internship)	None
Objectives of the Course	To teach chemical substances and chemical relations and chemical laws and concepts to students. To teach the students calculations about formulas and equations. To teach the students the structure of materials and materials and the changes in their structure. To be able to prepare the solutions that the students need and to teach the properties of the solutions. To ensure the student's usability in electrical-electronics engineering.
Contents of the Course	Chemistry, Thermodynamics, Reaction rate and equilibrium, Solubilisers, Solubility in water, Solubility in water, Solubility, Water solubility Equilibrium, Electrochemistry and corrosion, Periodic table and atomic structure, Chemical bond, redox reactions. Entalphy and Energy transformation

Learning Outcomes of Course

#	Learning Outcomes
1	Learn the aim of chemistry and fundamentals of matter.
2	Define chemical Factors
3	Learning SI system
4	Learn the periodic properties of elements

Course Syllabus

#	Subjects	Teaching Methods and Technics
1	The aim of chemistry, properties of matter, classification	Face to Face
2	The aim of chemistry, properties of matter, classification	Face to Face
3	SI system and logical numbers	Face to Face
4	SI system and logical numbers	Face to Face
5	Atoms and atom theories, orbital, elements, compounds	Face to Face
6	Atoms and atom theories, orbital, elements, compounds	Face to Face
7	Midterm	
8	Attractions between molecules, properties of liquids and solids	Face to Face
9	Attractions between molecules, properties of liquids and solids	Face to Face
10	Chemical reactions.	Face to Face

11	Chemical reactions.	Face to Face
12	Acid-base reactions, precipitation and solubility reactions, oxidation reduction reactions	Face to Face
13	Acid-base reactions, precipitation and solubility reactions, oxidation reduction reactions	Face to Face
14	Entalphy, Entrophy and Energy transformation	Face to Face
15	Revision	Face to Face
16	Final Exam	

Course Syllabus

#	Material / Resources	Information About Resources	Reference / Recommended Resources
1	Petrucci, Harwood, Herring Genel Kimya		

Method of Assessment

#	Weight	Work Type	Work Title
1	40%	Mid-Term Exam	Mid-Term Exam
2	60%	Final Exam	Final Exam

Relationship between Learning Outcomes of Course and Program Outcomes

#	Learning Outcomes	Program Outcomes	Method of Assessment
1	Learn the aim of chemistry and fundamentals of matter.	1,7	1,2
2	Define chemical Factors	7,11	1,2
3	Learning SI system	5,6,13	1,2
4	Learn the periodic properties of elements	3,5,7	1,2

PS. The numbers, which are shown in the column Method of Assessment, presents the methods shown in the previous table, titled as Method of Assessment.

Work Load Details

#	Type of Work	Quantity	Time (Hour)	Work Load
1	Course Duration	14	2	28
2	Course Duration Except Class (Preliminary Study, Enhancement)	0	0	0
3	Presentation and Seminar Preparation	0	0	0
4	Web Research, Library and Archival Work	1	2	2
5	Document/Information Listing	0	0	0
6	Workshop	0	0	0
7	Preparation for Midterm Exam	1	3	3
8	Midterm Exam	0	0	0
9	Quiz	0	0	0
10	Homework	5	6	30
11	Midterm Project	0	0	0
12	Midterm Exercise	0	0	0
13	Final Project	0	0	0
14	Final Exercise	0	0	0
15	Preparation for Final Exam	0	0	0
16	Final Exam	0	0	0
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