

# TOROS ÜNİVERSİTESİ

Vocational School  
Occupational Health And Safety

## Course Information

GENERAL CHEMISTRY					
Code	Semester	Theoretical	Practice	National Credit	ECTS Credit
		Hour / Week			
İSG113	Fall	2	0	2	3

<b>Prerequisites and co-requisites</b>	
<b>Language of instruction</b>	Turkish
<b>Type</b>	Required
<b>Level of Course</b>	Associate
<b>Lecturer</b>	Lect. Dilşat AKGÜL
<b>Mode of Delivery</b>	Face to Face
<b>Suggested Subject</b>	
<b>Professional practise ( internship )</b>	None
<b>Objectives of the Course</b>	Aim of this course is to inform students about structure of atom, molecules, compounds and chemical reaction types, state of matter, chemical bonds and molecular shapes.
<b>Contents of the Course</b>	Matter and elements. Compounds and nomenclature of compounds. Mixtures. Measurements and mole concept. Atomic structure and periodic table. Chemical reactions and reaction stoichiometry. Chemical bonds, chemical bonding theories. Intermolecular forces and liquids, solids, gases. Solutions, concentration units, solutions and their properties. Entrophy and Entalphy .

## Learning Outcomes of Course

#	Learning Outcomes
1	Learn the aim of chemistry and fundamentals of matter.
2	Chemical Factors
3	Learning SI system
4	Learn the periodic properties of elements

## Course Syllabus

#	Subjects	Teaching Methods and Technics
1	The aim of chemistry, properties of matter, classification	Lecture
2	The aim of chemistry, properties of matter, classification	Lecture
3	SI system and logical numbers	Lecture
4	SI system and logical numbers	Lecture
5	Atoms and atom theories, orbital, elements, compounds	Lecture
6	Atoms and atom theories, orbital, elements, compounds	Lecture
7	Midterm	
8	Attractions between molecules, properties of liquids and solids	Lecture
9	Attractions between molecules, properties of liquids and solids	Lecture
10	Chemical reactions.	Lecture
11	Chemical reactions.	Lecture

12	Acid-base reactions, precipitation and solubility reactions, oxidation reduction reactions	Lecture
13	Acid-base reactions, precipitation and solubility reactions, oxidation reduction reactions	Lecture
14	Revision	Lecture
15	Entalphy, Entropy and Energy transformations	Lecture
16	Final Exam	

### Course Syllabus

#	Material / Resources	Information About Resources	Reference / Recommended Resources
1	Petrucci, Harwood, Herring Genel Kimya		

### Method of Assessment

#	Weight	Work Type	Work Title
1	40%	Mid-Term Exam	Mid-Term Exam
2	60%	Final Exam	Final Exam

### Relationship between Learning Outcomes of Course and Program Outcomes

#	Learning Outcomes	Program Outcomes	Method of Assessment
1	Learn the aim of chemistry and fundamentals of matter.	1,7	1,2
2	Chemical Factors	7,11	1,2
3	Learning SI system	5,6,13	1,2
4	Learn the periodic properties of elements	3,5,7	1,2

PS. The numbers, which are shown in the column Method of Assessment, presents the methods shown in the previous table, titled as Method of Assessment.

### Work Load Details

#	Type of Work	Quantity	Time (Hour)	Work Load
1	Course Duration	14	2	28
2	Course Duration Except Class (Preliminary Study, Enhancement)	0	0	0
3	Presentation and Seminar Preparation	0	0	0
4	Web Research, Library and Archival Work	1	2	2
5	Document/Information Listing	0	0	0
6	Workshop	0	0	0
7	Preparation for Midterm Exam	0	0	0
8	Midterm Exam	0	0	0
9	Quiz	0	0	0
10	Homework	5	6	30
11	Midterm Project	5	6	30
12	Midterm Exercise	0	0	0
13	Final Project	0	0	0
14	Final Exercise	0	0	0
15	Preparation for Final Exam	0	0	0
16	Final Exam	0	0	0
				<b>90</b>